

# **Calibration of a PHREEQC Based Geochemical Model to Predict Surface Water Discharge Compositions from an Operating Uranium Mill in the Athabasca Basin**

**John J. Mahoney, Ph.D.**

**Mahoney Geochemical Consulting, LLC**

**Ryan A. Frey**

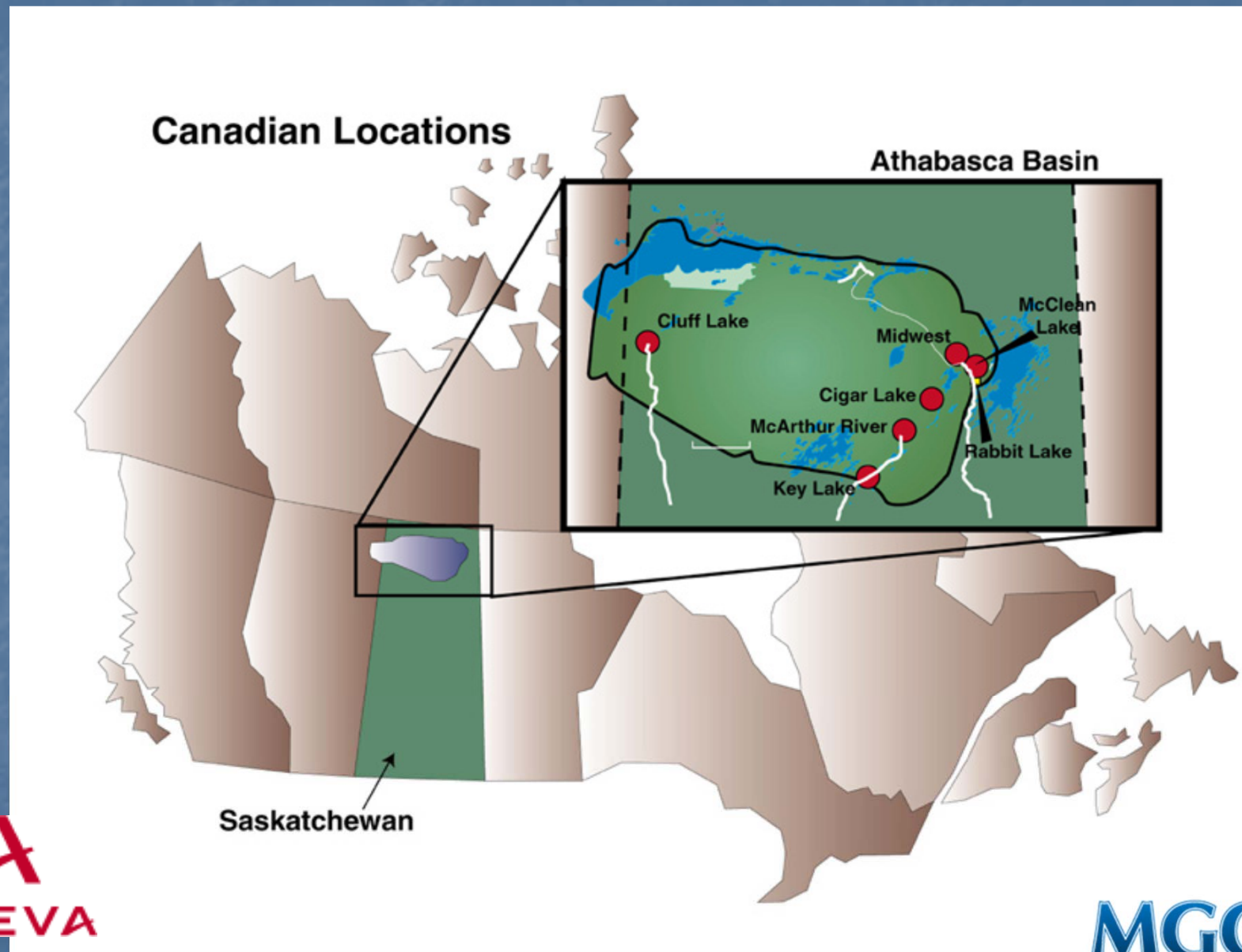
**AREVA Resources Canada, Inc.**



# Objectives

- **Develop predictive model to estimate concentrations in the Sink Vulture Treated Effluent Management System (SVTEMS) for AREVA Resources Canada McClean Lake Mill**
  - Sink Reservoir, Vulture and McClean Lakes
  - PHREEQC based calculations for geochemistry
  - Employ PHREEPLOT for data fittings
- **Model designed to predict concentrations in response to changing conditions, including:**
  - Different ores
  - Different processes
  - Different waters sources
  - Changing treatment conditions
- **This is a batch mixing model**
  - Think well mixed beakers
  - Each model represents one year
  - No year-to-year carry over in models

# Location Map



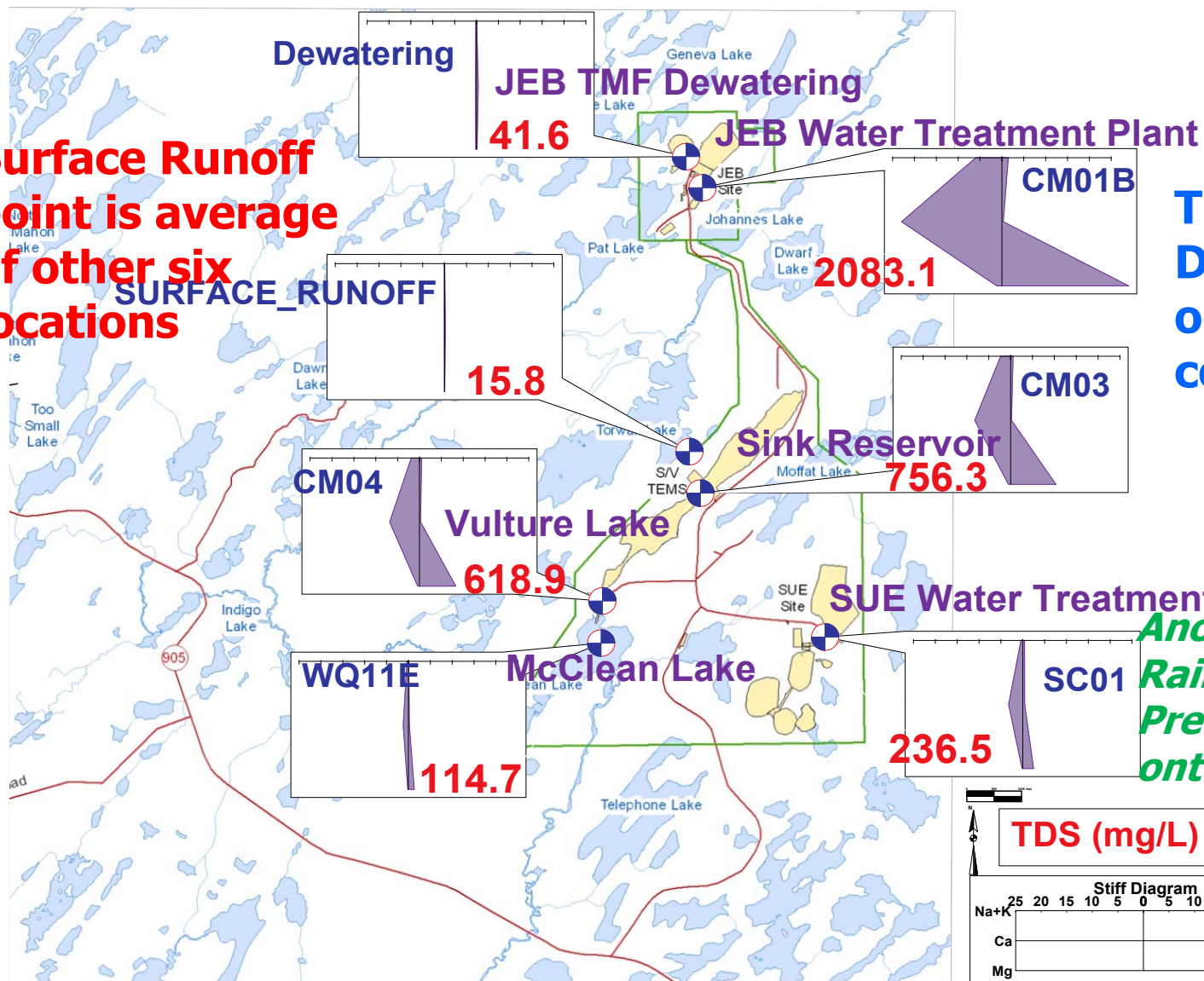


# McClean Lake Mill and JEB TMF





**Surface Runoff  
point is average  
of other six  
locations**

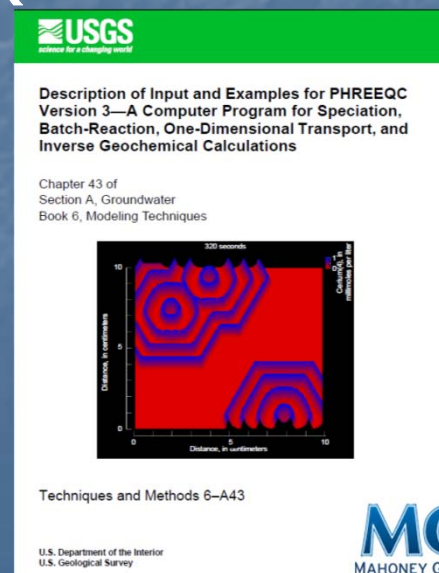


**TDS and Stiff  
Diagrams based  
on average  
concentrations**

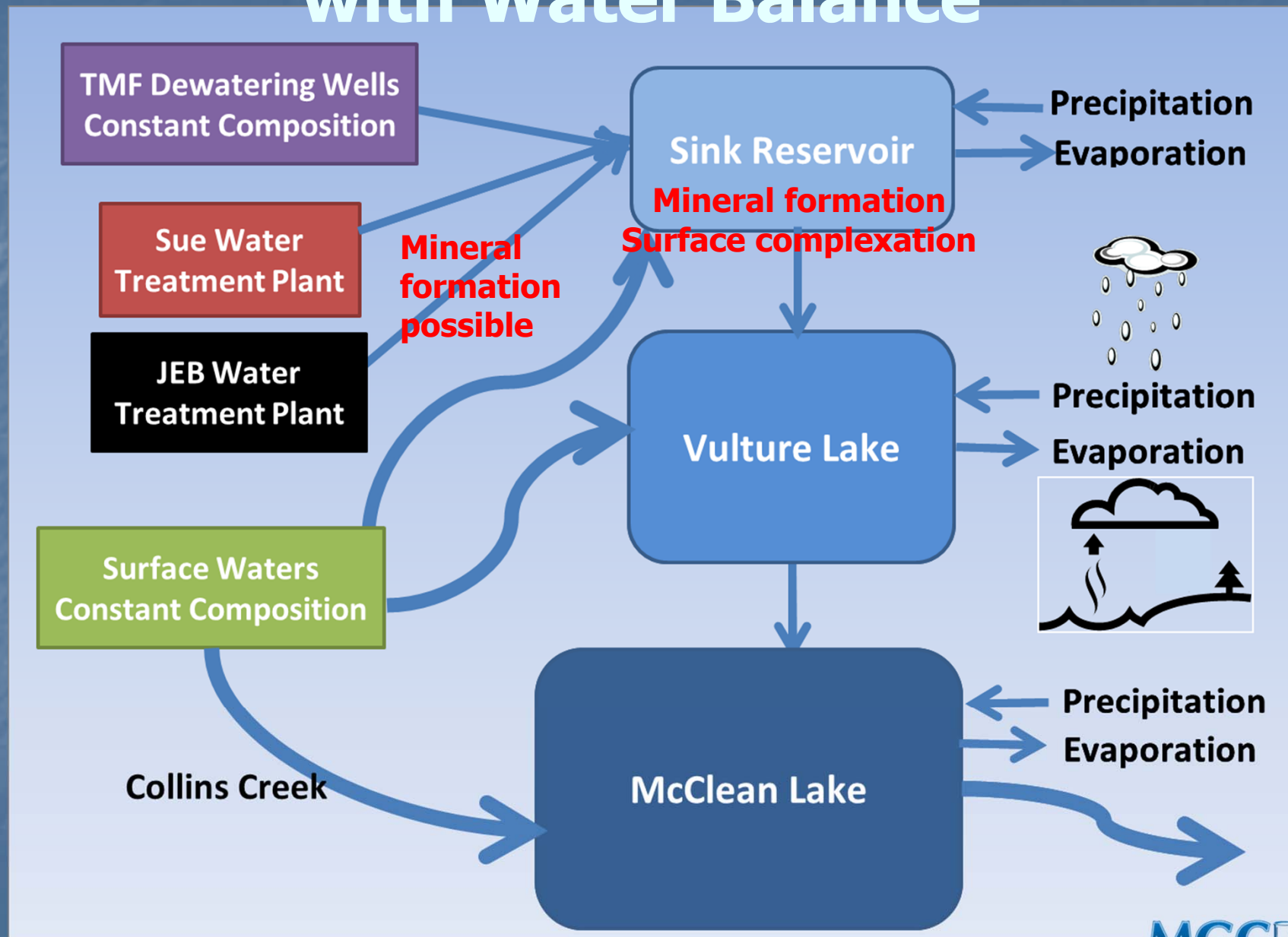
*Another water source  
Rain or snow  
Precipitation directly  
onto lake surfaces*

# What is PHREEQC?

- PHREEQE (1980) pH Redox Equilibrium Equations
- Performs aqueous geochemical modeling
  - Calculates activities of a component set
  - Calculates distribution of species in solution
  - Calculates mineral saturation indices
  - Mass transfer capabilities - mix, equilibrate, precipitate
- PHREEQC (1995, 1999, 2013) written in C language
  - Batch, Interactive and COM versions available
  - Public domain (USGS) ongoing support (David Parkhurst)
- PHREEQC Capabilities
  - Mixing of multiple solutions
  - Equilibration with gases/minerals
  - Surface complexation models
  - Ion exchange
  - Irreversible reactions/evaporation
  - Kinetics – Rate controlled processes



# Conceptual Hydrogeochemical Model with Water Balance





# Updated Thermo Database

- **WATEQ4F.DAT PHREEQC Version 2.0 HEAVILY REVISED**
- **Uranium**
  - OECD NEA (Guillaumont 2003) for uranyl complexes
  - Diffuse layer parameters for HFO uranyl surface complexation (Mahoney et al. 2009) for NEA database replaced Dzombak and Morel (D&M) values
  - Dong and Brooks (2006) - divalent uranyl carbonates
  - Removed  $\text{UO}_2(\text{HPO}_4)_2^{-2}$  and  $\text{UO}_2(\text{H}_2\text{PO}_4)_3^-$  per Grenthe et al. 1992
- **Arsenic**
  - Langmuir, Mahoney & Rowson (2006) for metal As complexes, and Scorodite
  - Marini and Accornero (2007) more metal As complexes, with 2010 corrections
  - Diffuse layer surface complexation constants from Gustafsson and Bhattacharya (2007) replaced D&M values
- **Molybdenum**
  - Surface reactions/values by Gustafsson (2003) replaced D&M values
  - $\text{NiMoO}_4$  solid by Morishita and Navrotsky (2003)
- **Gibbsite Surface Complexes – Karamalidis and Dzombak (2010)**

# Overview

- Use Water Balance from 2000 to 2011 to estimate proportions of different waters
- Water compositions from monitoring points
- Originally used EXCEL to:
  - Calculate mixing proportions for PHREEQC input
  - Calculate evaporation losses
  - Write majority of input file
  - Block and copy into PHREEQC

## Water Balance Calculations (m<sup>3</sup>/year)

	Direct Sources			Precipitation Rain/Snow			Calc groundwater, streams & runoff			Evaporation		
Year	Dewatering Wells	CM01B	SC01	Sink	Vulture	McClean	Sink	Vulture	McClean	Sink	Vulture	McClean
2000	2110337	1548498	1075486	89842	269525	822875	0	1773184	42029593	85750	257250	785400
2001	2328796	1476709	1303234	108731	326193	995887	0	1987207	40740609	85750	257250	785400
2002	1314516	1300817	263083	163146	489437	1494280	0	900954	64639280	85750	257250	785400
2003	1590610	1561785	59386	164959	494876	1510885	0	1300923	82451780	85750	257250	785400
2004	1619357	1559263	979574	84795	254384	776648	0	1541937	34705257	85750	257250	785400
2005	1460828	1635128	2537086	187499	562496	1717333	0	1823447	66190958	85750	257250	785400

# Water Balance Converted to Mixing Proportions

units m <sup>3</sup> /year		Year	Solution #	2000	2001	2002
SINK INPUTS		D-Wells to Sink	200	2,110,337	2,328,796	1,314,516
	001	CM01B	102000001	1,548,498	1,476,709	1,300,817
		PRECIPITATION	100	89,842	108,731	163,146
	002	SC01	2000002	1,075,486	1,303,234	263,083
		Calc groundwater, streams & runoff	300	0	0	0
		EXTRA SOURCE NOT INCLUDED	400	0	0	0
	003	WATER IN SINK AT START OF YEAR	2000003	395000	395000	395000
VOLUME Sum use for mixing proportions				5,219,163	5,612,470	3,436,562
Calc input to Sink FROM WATER BALANCE				4,738,413	5,131,720	2,955,812
Evaporation LOSS			Sink	85,750	85,750	85,750
		CHECK CALCULATION		5,133,413	5,526,720	3,350,812
MIXING CALCULATION		D-Wells to Sink		0.404	0.415	0.383
		CM01B		0.297	0.263	
		PRECIPITATION		0.017	0.019	
		SC01		0.206	0.232	
		Calc groundwater, streams & runoff	300	0.000	0.000	
		EXTRA SOURCE NOT INCLUDED		0.000	0.000	
		SINK WATER PREVIOUS YEAR		0.076	0.070	
		SUM		1.000	1.000	
EVAPORATION CALCULATION		FRACTION EVAP LOSS		0.0164	0.0153	0.0250
		Reactant Amount for evaporation		-0.913	-0.849	-1.386
		MIXING PROPORTION FOR volume correction		1.016	1.015	1.025

Can easily expand and  
add additional years  
Extra water (blank at this  
time) available to modify  
model







- ***"Pluralitas non est ponenda sine neccesitate"***
  - Pluralities should not be posited without necessity
  - One should not increase, beyond what is necessary, the number of entities required to explain anything
- **Using generally accepted geochemical principles with only slight adjustments**
  - Dominantly mixing processes
  - Equilibration with atmospheric gases and simple minerals
  - Adjusted one solubility product constant (Powellite)
  - Lower surface site density values - about 1/2

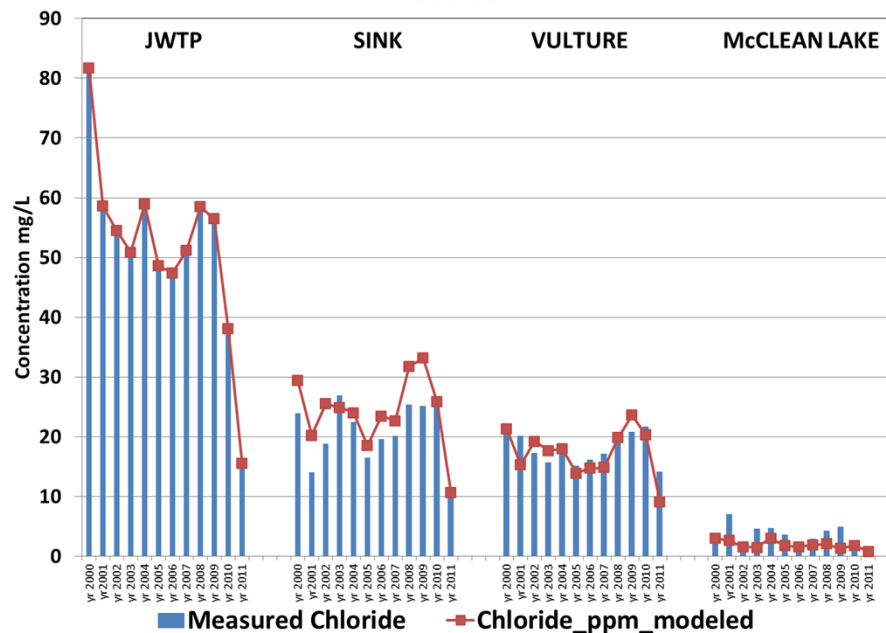


# Conservative Species

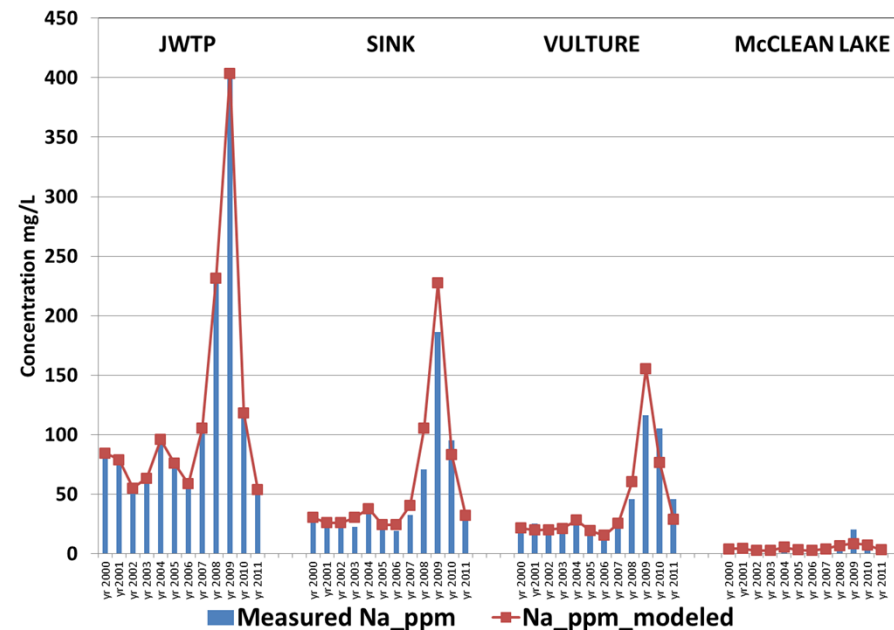
- **Na, K, Mg, Cl - Most conserved**
  - Good matches between observed and modeled values
  - Mixing based upon water balance – only process
- **Minimal attenuation - Sulfate, Ca**
  - Gypsum ( $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ ) precipitation if oversaturated
  - Only three samples from JWTP oversaturated
- **Aluminum appears to be conserved**
  - Most samples
  - Early models looked at gibbsite [ $\text{Al}(\text{OH})_3$ ] or boehmite ( $\text{AlOOH}$ ) precipitation
  - Produced poorer fits downstream
  - Maybe sulfate inhibits Al precipitation?



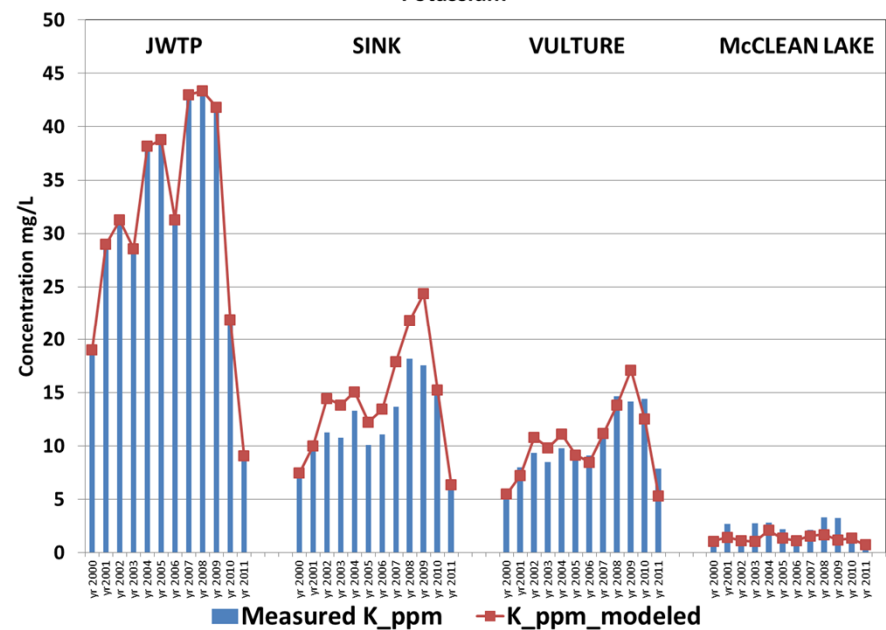
### Chloride



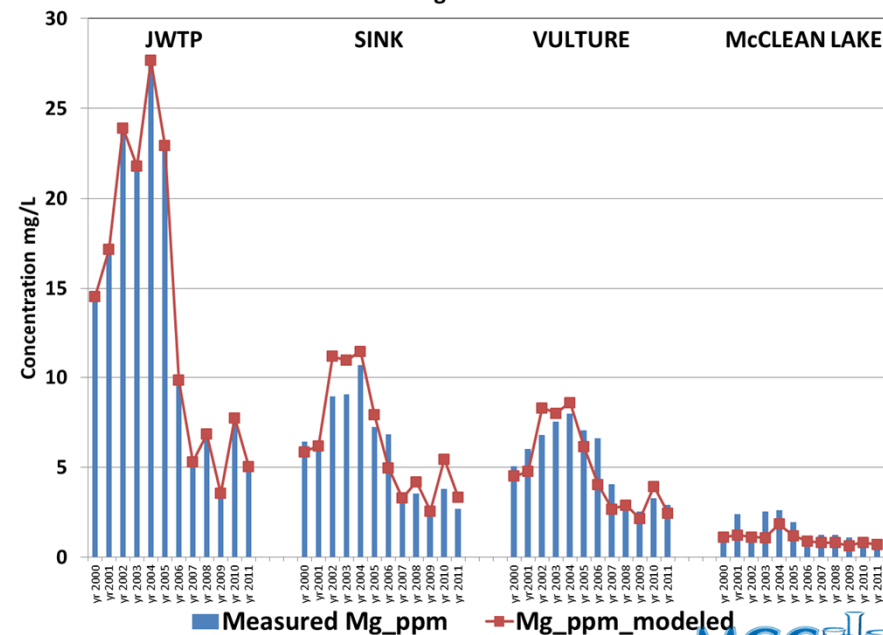
### Sodium



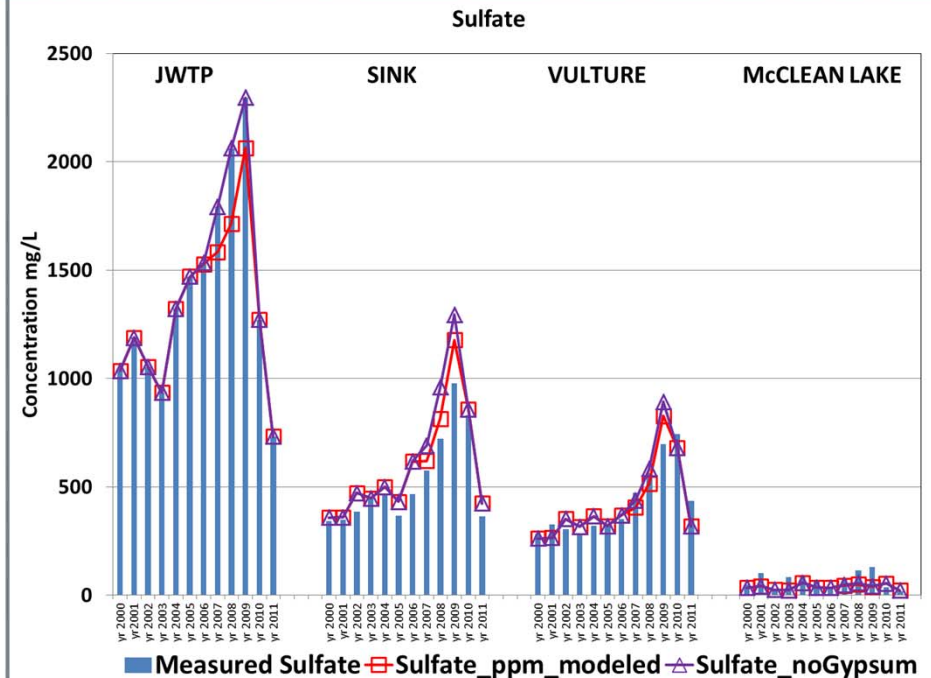
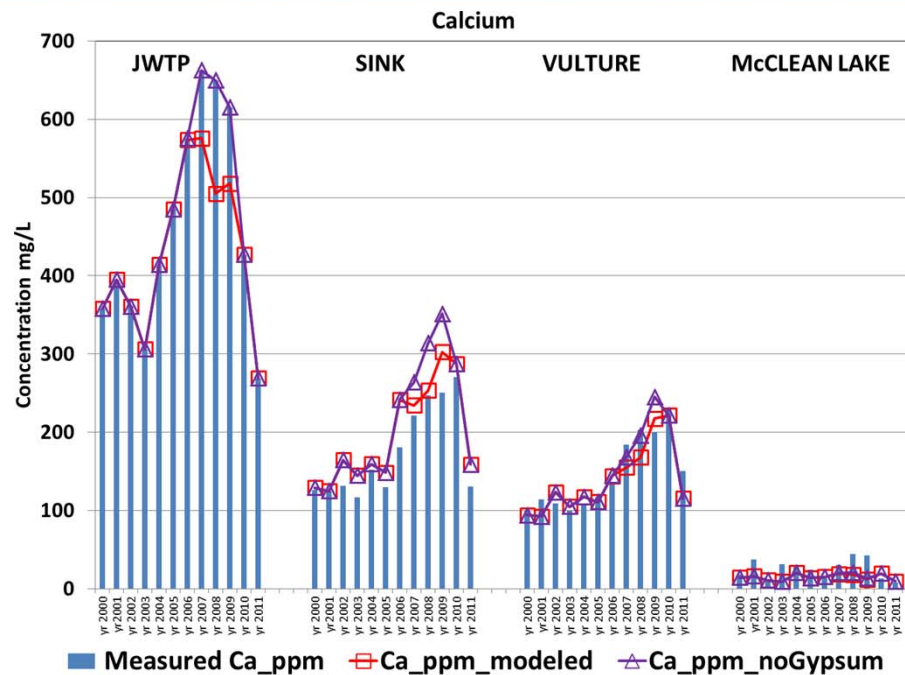
### Potassium



### Magnesium

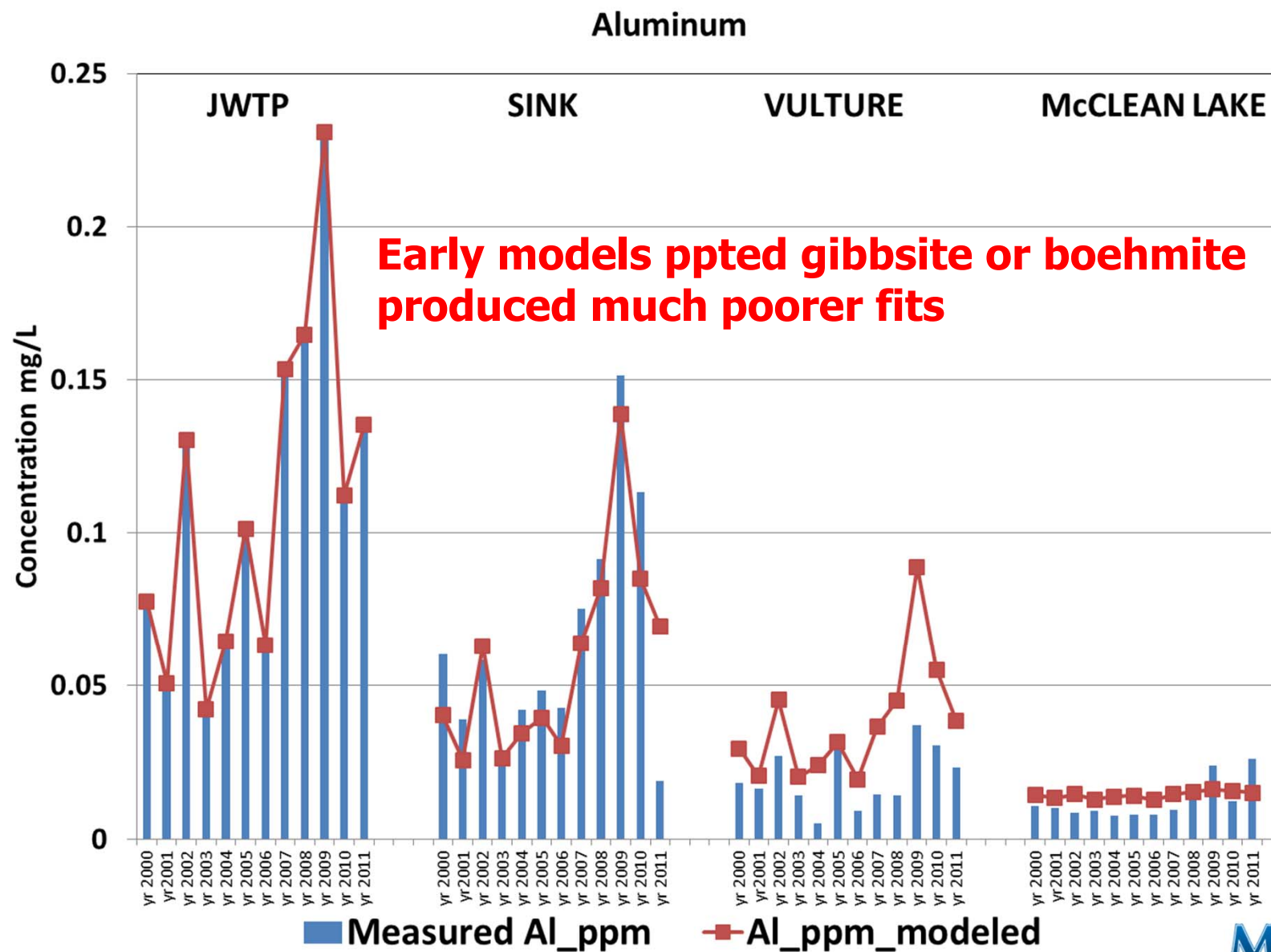


# Calcium and Sulfate



**Gypsum Precipitation in Three Waters**

# Aluminum

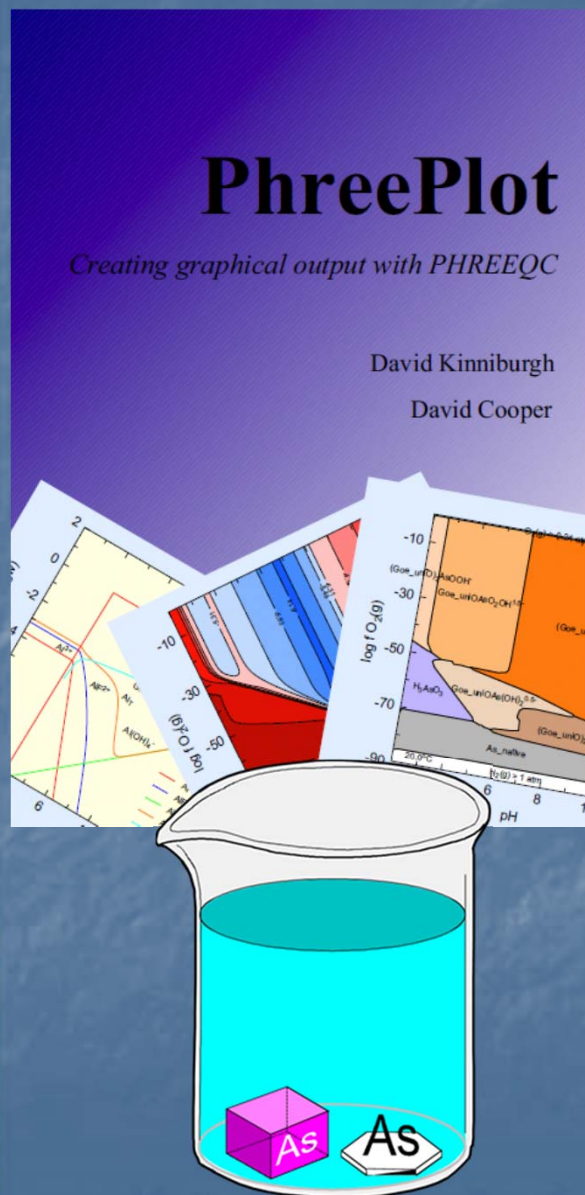




# Reliability of Water Balance and Analyses

- Agreement between modeled and observed concentrations for conserved species demonstrate that the conceptual model, water balance, analyses and the calculations using PHREEQC are correct
- Allows insights into reactive trace metals and other species such as ammonia

# Optimization with PhreePlot



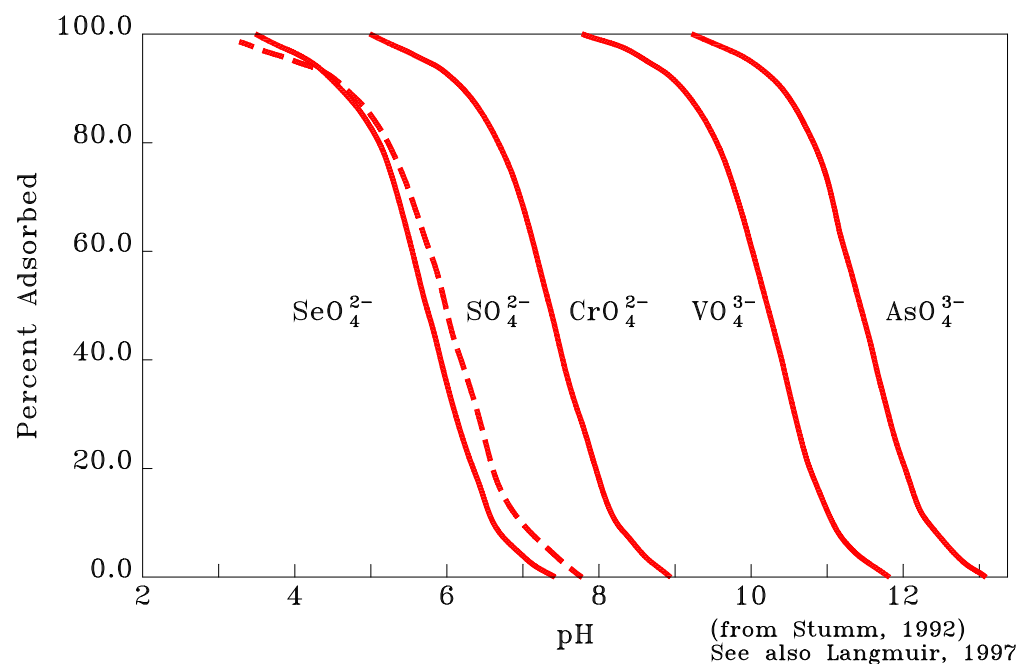
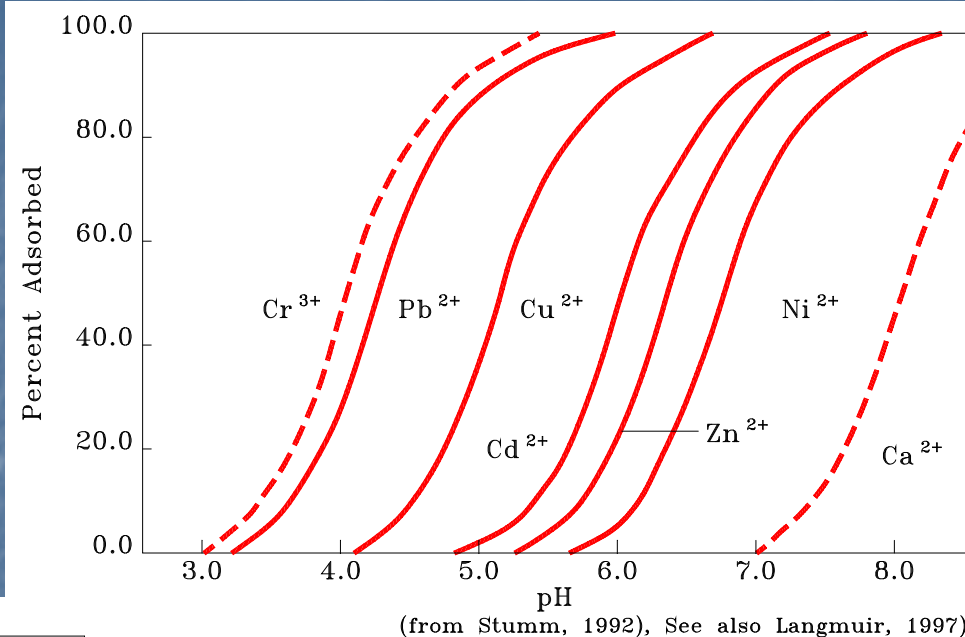
- PHREEQC with data fitting functions
  - Iterative process between observations and modeled concentrations
  - Calculates Weighted Residual Sum of Squares
  - Minimizes (W)RSS
  - Rigorously fitted parameters
- Used to model As and U surface complexation on Ferrihydrite (HFO)
  - Sorption onto a surface
  - Hydrated Ferric Oxide (HFO)
  - HFO comes from dissolved iron in discharges
- Originally adjusted HFO concentrations – JWTP, then Sink Reservoir
  - Applied to estimate HFO sites\_per\_mole value based upon HFO precipitation
- Fit molybdenum solubility product

# Surface Complexation Reactions

Provides alternative method to remove compounds from solution

Often occurs at concentrations less than the solubility limit

Anions decrease sorption on HFO with increasing pH



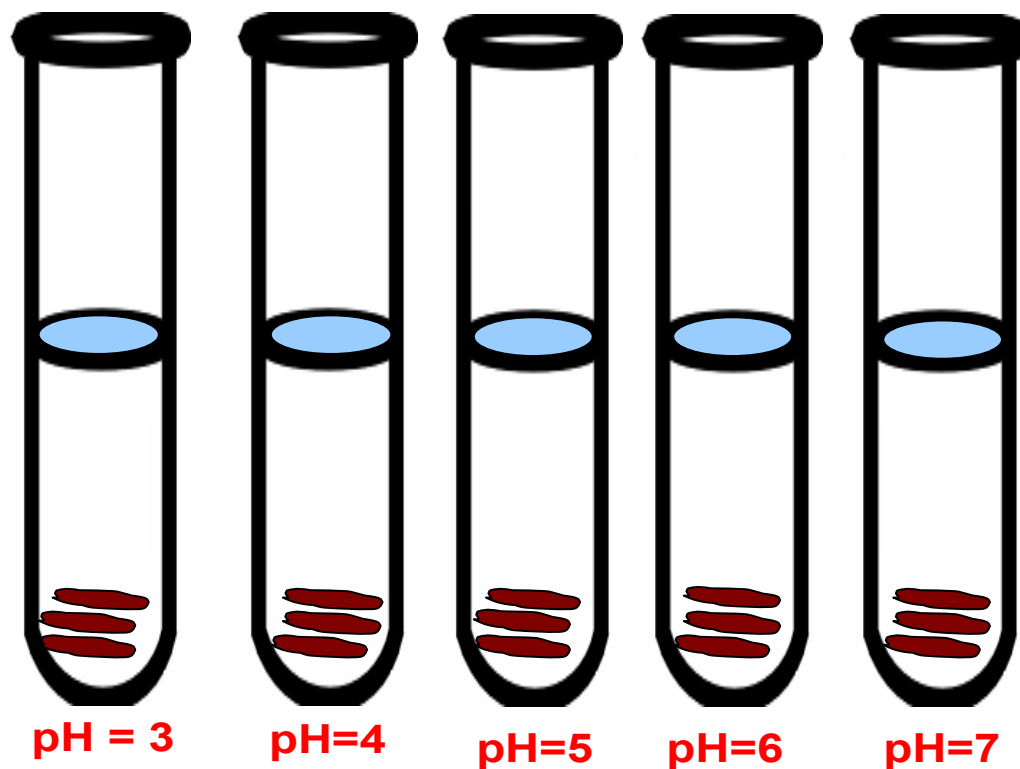
Cations increase sorption on HFO with increasing pH

**Because of complexation with carbonate uranium increases then decreases**



# Uranium Adsorption Experiments

Initial Composition U(VI) = 2,380  $\mu\text{g/L}$ , in 0.1  $\text{NaNO}_3$

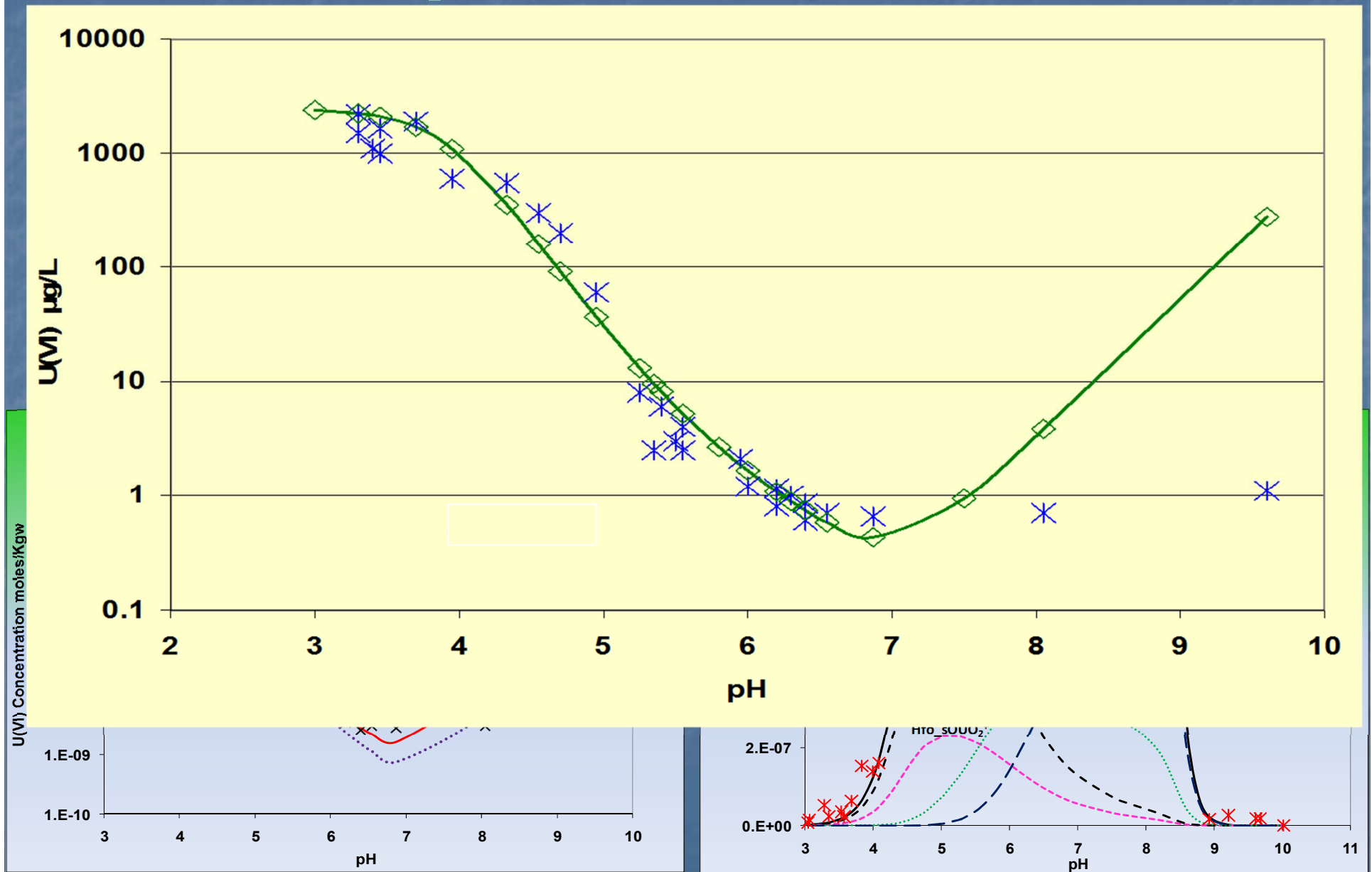


1g/L  
Ferrihydrite

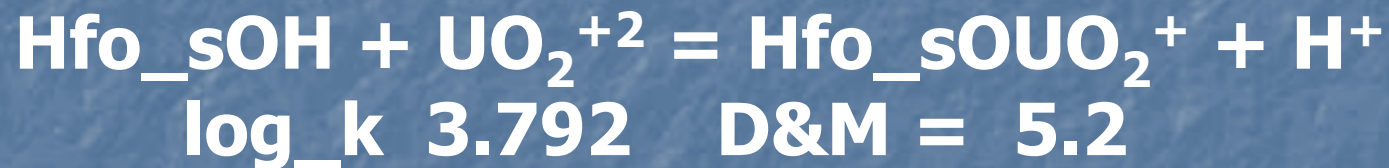
2,380  $\mu\text{g/L}$  =  $1.0 \times 10^{-5}$  m U

**No Carbonate Present**

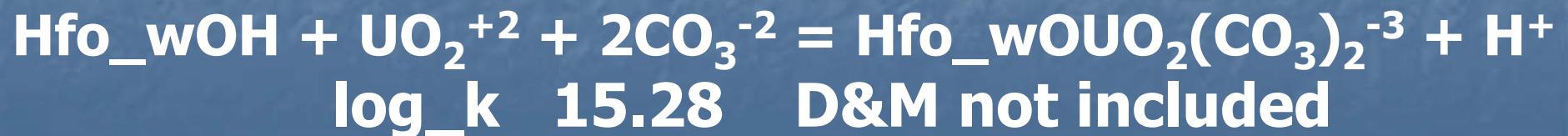
# Refit of Uranium Surface Complexation Reactions



# Surface Complexation Reactions for NEA database



**Need to estimate surface site concentrations of  
Hfo\_wOH and Hfo\_sOH**

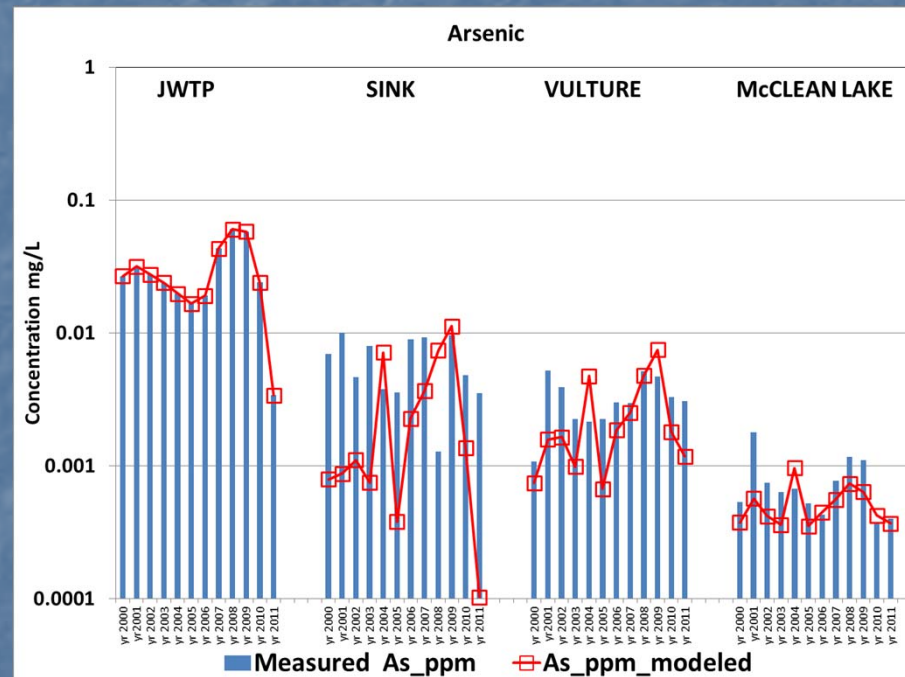
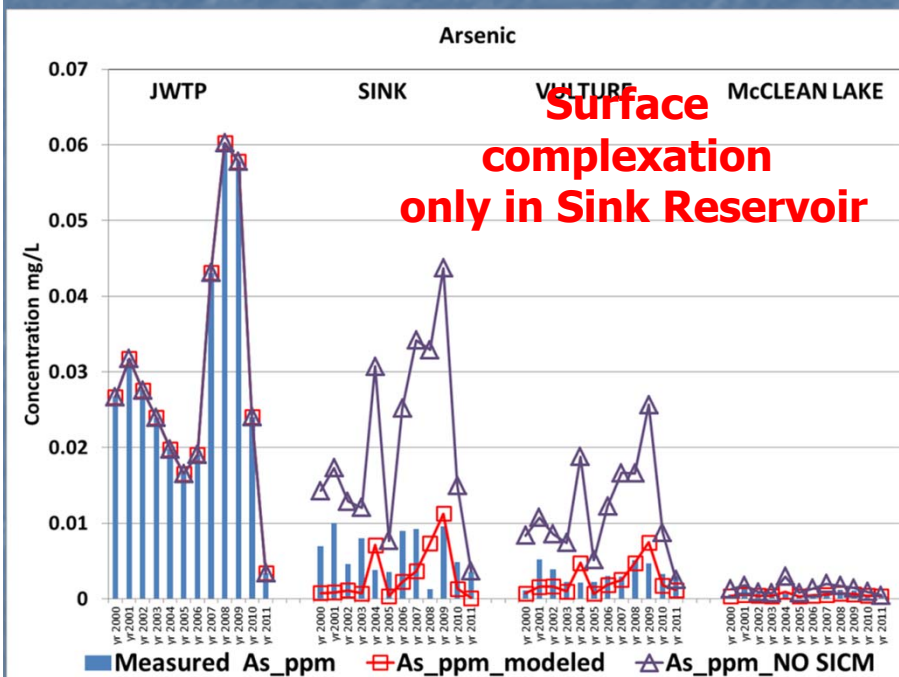




# Arsenic With and Without Surface Complexation

Linear Concentration Scale

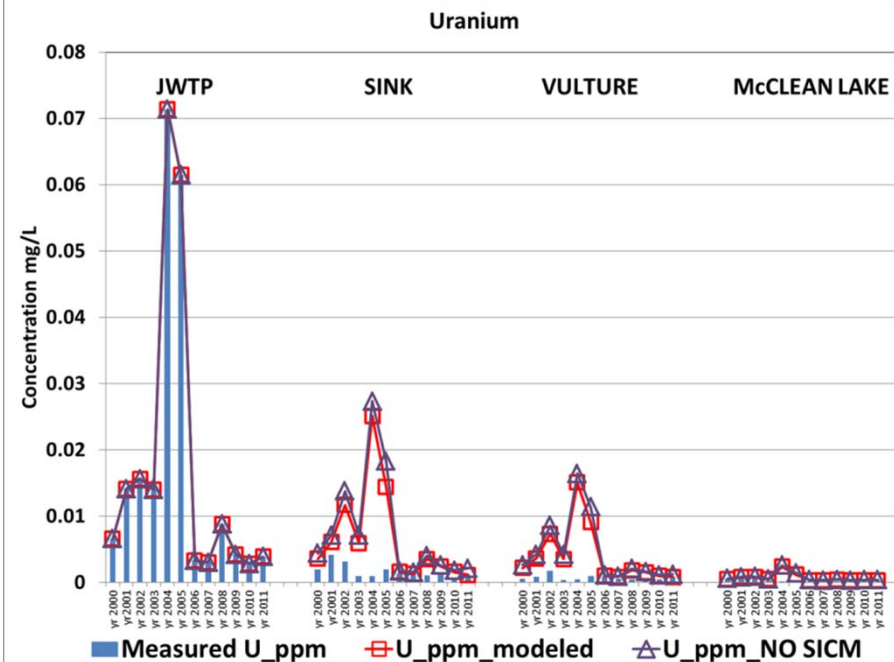
Log Concentration Scale



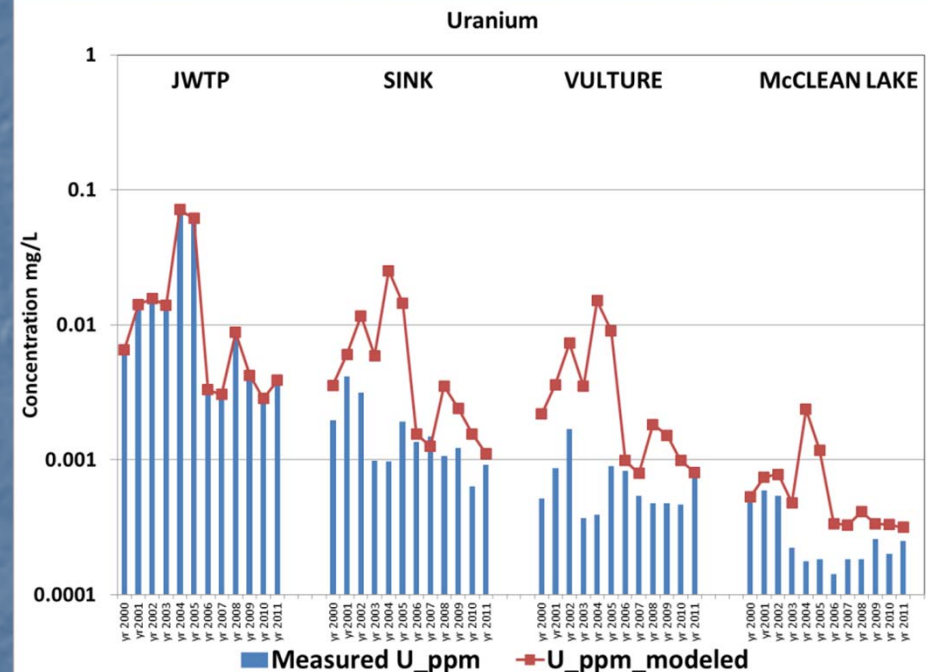
**Final fit from PHREEPLOT**  
**Uses 0.09 for first six points, and**  
**0.129 for last six points for site density function**

# Uranium

## Linear Concentration Scale



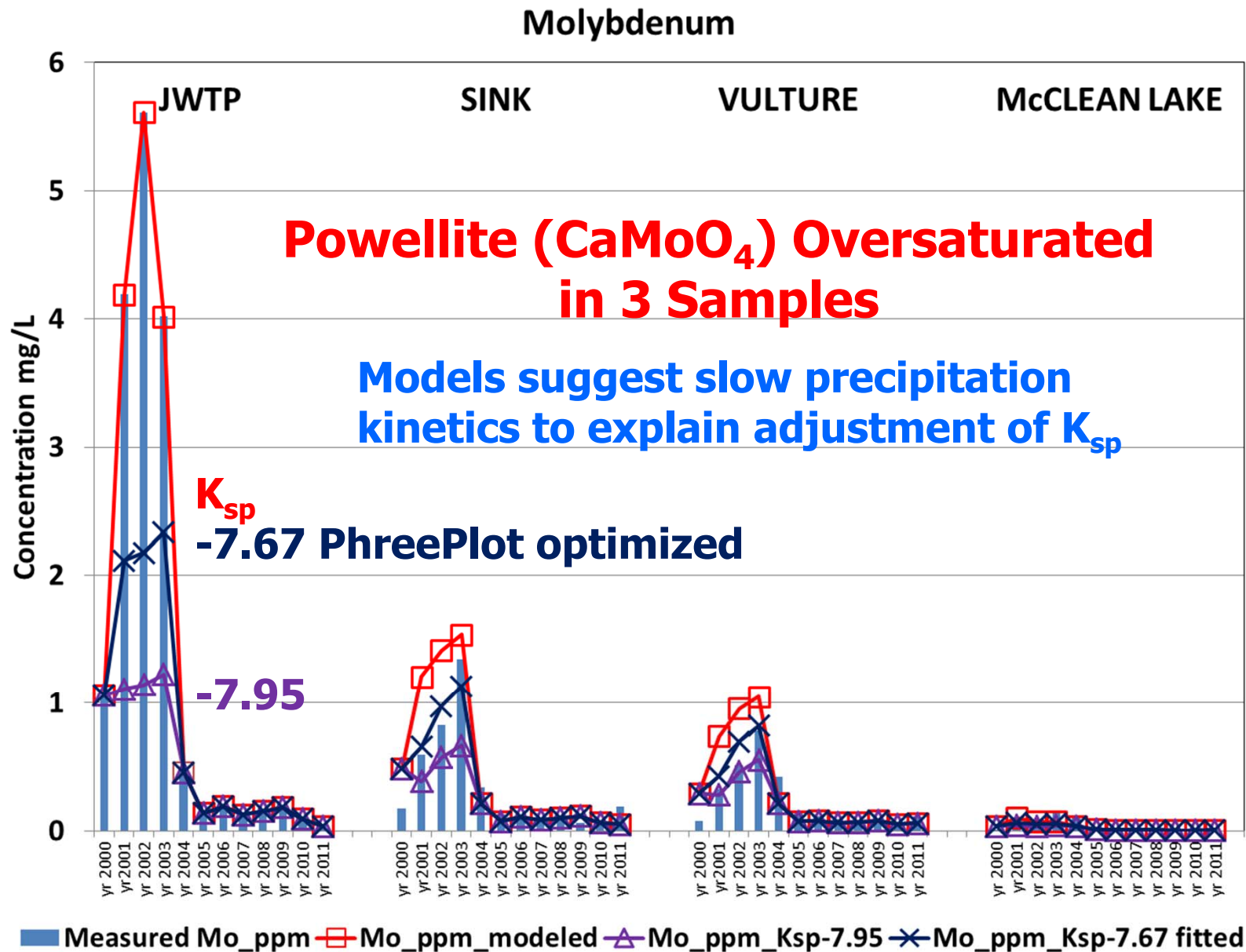
## Log Concentration Scale



Deliberately selected lower site concentration to match As  
More conservative setup for Uranium  
i.e., produces higher modelled values

# Molybdenum Precipitation - Powellite

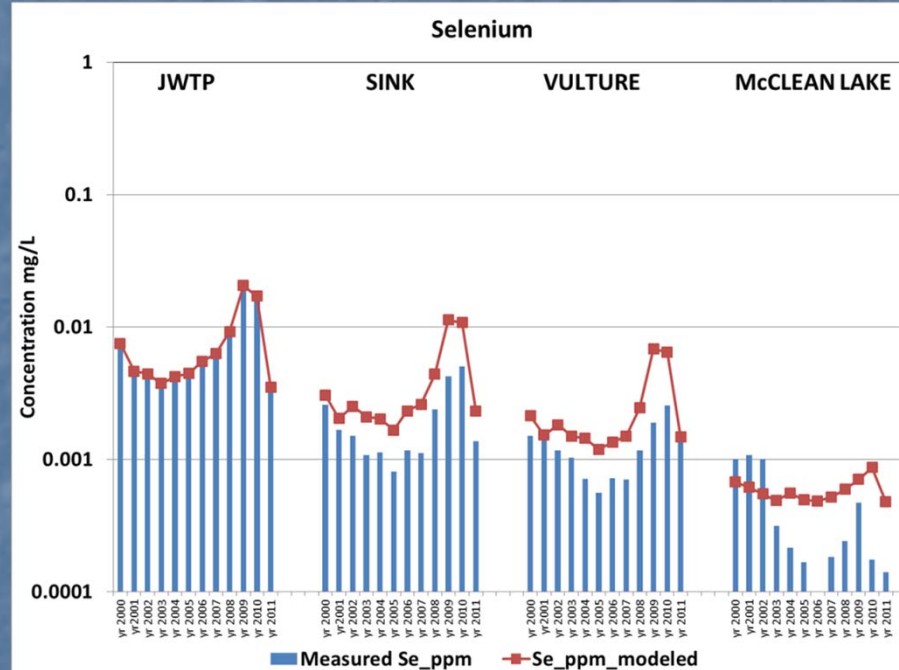
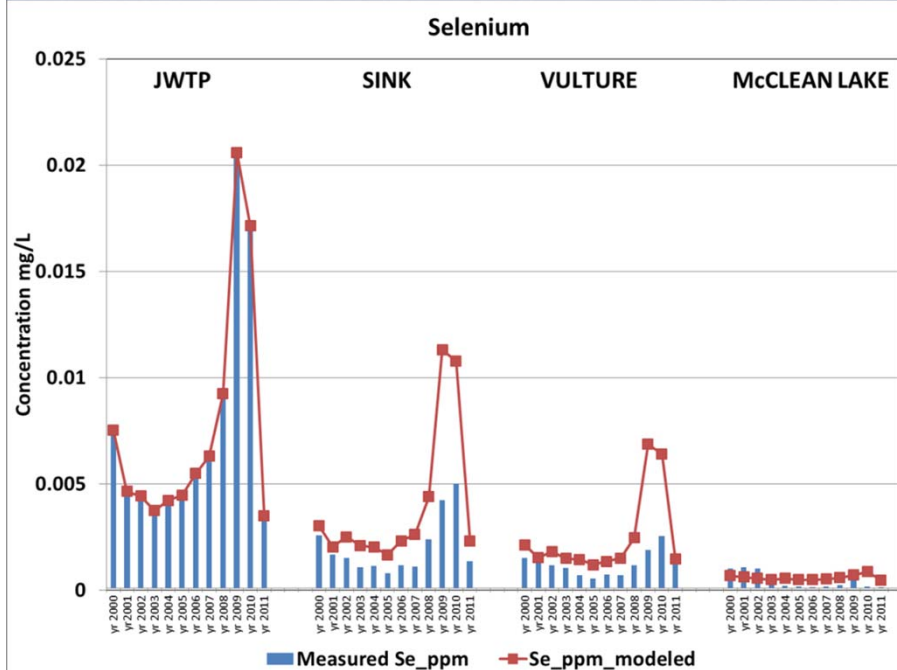
$K_{sp} = -7.95$  and  $-7.67$  (PhreePlot Fitted)





# Selenium

Linear Concentration Scale      Log Concentration Scale



# Continuing Work

- **Evaluate Seasonal and Redox Effects**
  - Three samples per year
  - Evaluate seasonal effects
  - Better handle on redox processes
  - Iron related processes
- **pH and better estimate of CO<sub>2</sub>(g) PP**
- **Iron needs additional work**
  - Poor observed/modeled matches in Sink Reservoir, but
  - Good matches in Vulture and McClean Lakes
- **Possible role of humic substances**
  - Redox control
  - Sorbent phase

# Conclusions

- **Good predictions for conservative species:**
  - Cl, Mg, K, Na, - Ca, Sulfate, - (Al)
  - Reliability of water balance and analyses
  - Mainly mixing of waters and dilution
  - No reason to change water balance/mixing proportions
- **Better than order of magnitude fits for trace metals**
  - Often within a factor of  $\pm 2.0$ , but most within 25% RPD
  - Some elements agree down to ppb levels
- **Surface complexation attenuates As and U**
- **Precipitation of powellite ( $\text{CaMoO}_4$ ) explains Mo concentrations**
- **Conservative assumptions used in model setup**
  - Not all possible attenuation processes included
  - Slight overestimates for uranium, selenium, ammonia

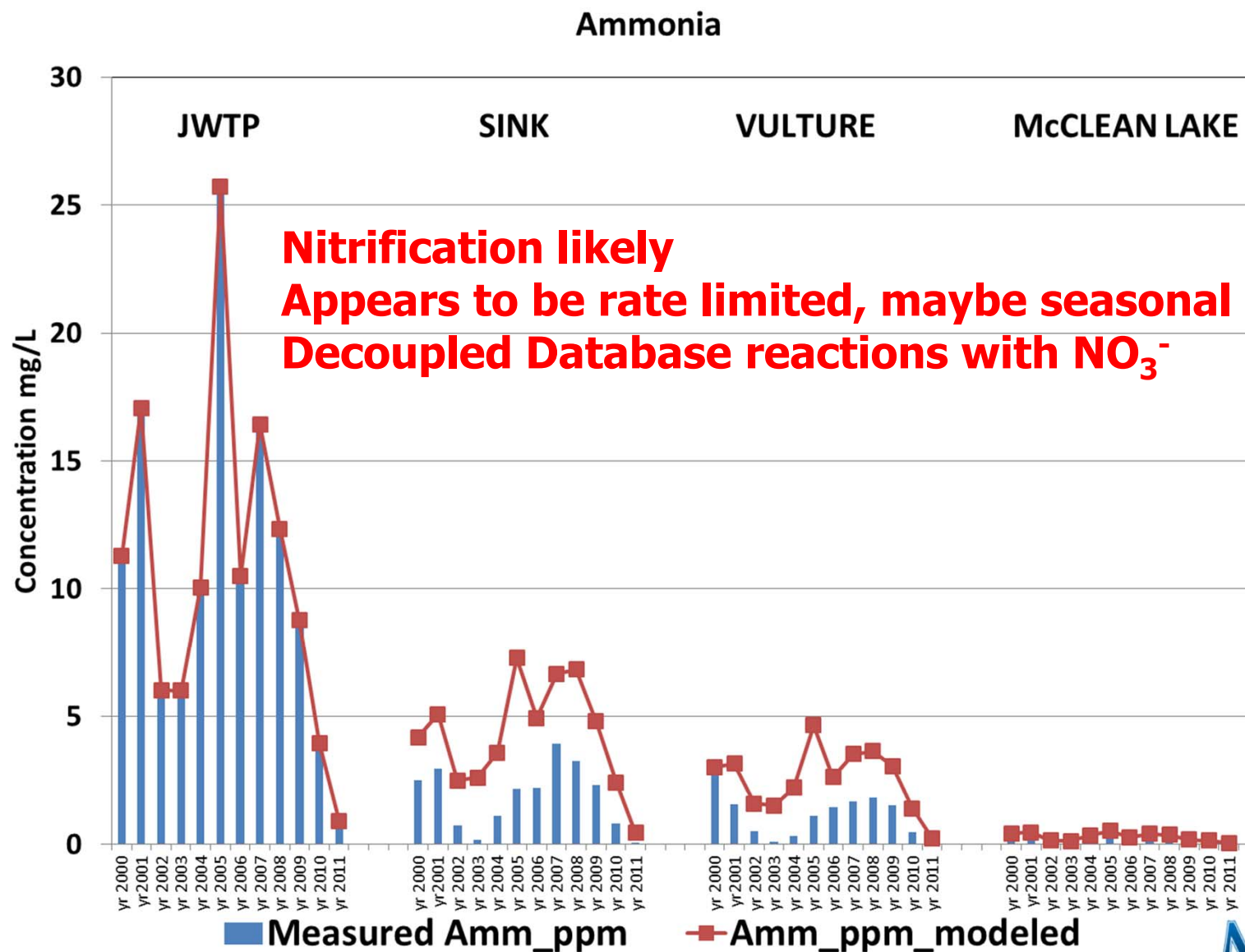




**Questions?**

An aerial photograph showing a large industrial complex, likely a mining or processing plant, situated in a vast, open landscape. The facility includes several large buildings, storage tanks, and a network of roads. In the foreground, a large, circular, open-pit mine is visible, filled with a greenish liquid. The surrounding area is characterized by dense green forests and numerous small lakes or ponds. The sky is clear and blue.

# Ammonia is Consumed





# SURFACE SITE CALCULATIONS

**SURFACE 4** # user defined surface site concentration

**Hfo\_wOH** <Hfo\_CONC> **600** **0.045**

**Hfo\_SOH** <strong>

Alternative method EQ\_P and Surface keywords linked

**EQUILIBRIUM\_PHASES 3**

**O2(g) -0.7 10** # oxygen at atmospheric pressure

**CO2(g) -3.5 10** # carbon dioxide at atmospheric pressure

**SURFACE 3**

**Hfo\_wOH** Ferrihydrite equilibrium\_phase <sitedens> 53400

**Hfo\_sOH** Ferrihydrite equilibrium\_phase <strong>

Uses PHREEPLOT to adjust <sitedens> to optimize fit between measured and modeled arsenic concentration in Sink, Vulture and McClean Lakes

PHREEPLOT will also calculate <strong> site concentration in agreement with the <sitedens> value and original D&M assumptions



# IPhreeqcCOM Module

- **Original Version of Model**
  - EXCEL File Setup
  - model is copied into PHREEQC
  - Water compositions pre-loaded
  - Supporting keywords pre-loaded
  - Output file and selected output file
- **Now IphreeqcCOM**
  - Component Object Module
  - Can run as module in VBA macro in EXCEL
  - Whole process is run within EXCEL
  - Water balance, mixing calculations, PHREEQC syntax,
  - Run macro
  - Reads PHREEQC input file
  - Write selected output to sheet
  - Facilitates preparation of graphics
  - Improves tracking of models

